



Units 3 and 4 Chemistry

Practice Exam Solutions

Stop!

Don't look at these solutions until you have attempted the exam.

Any questions?

Check the Engage website for updated solutions, then email practiceexams@ee.org.au.

Section A – Multiple-choice questions

Question 1

The correct answer is D.

Question 2

The correct answer is D.

Question 3

The correct answer is D.

Question 4

The correct answer is D.

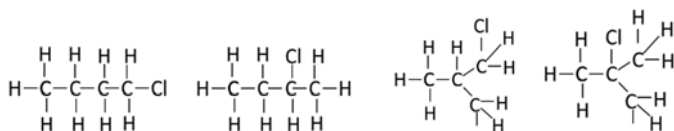
Question 5

The correct answer is C.

Question 6

The correct answer is D.

The following are all possible structural isomers:



Question 7

The correct answer is B.

Question 8

The correct answer is A.

Question 9

The correct answer is A.

Question 10

The correct answer is D.

Question 11

The correct answer is C.

Question 12

The correct answer is D.

Question 13

The correct answer is C.

Question 14

The correct answer is B.

Question 15

The correct answer is C.

Question 16

The correct answer is C.

Changing the atmospheric pressure would alter the rate of a reaction in the gaseous phase but not for a reaction involving a solid and a liquid.

Question 17

The correct answer is A.

Question 18

The correct answer is D.

Question 19

The correct answer is C.

As the reaction is exothermic ($\Delta H < 0$), a decrease in temperature will favour the forward reaction, thus decreasing the amount of ClF_3 in the mixture and increasing the amount of F_2 and Cl_2 .

Question 20

The correct answer is C.

Question 21

The correct answer is C.

If the coefficients in a balanced equation are multiplied by a factor, n , the equilibrium expression is raised to the n th power.

The equilibrium constant for the reverse reaction is the reciprocal ($1/K$) of the equilibrium constant for the forward reaction.

So in the first reaction, where the coefficients are doubled, the new equilibrium constant is the square of the original.

In the second reaction, where the coefficients are halved, you must find the square root of the original equilibrium constant, and as the reaction is also reversed, the reciprocal of this number is the final equilibrium constant.

Question 22

The correct answer is B.

Question 23

The correct answer is A.

Question 24

The correct answer is B.

Question 25

The correct answer is A.

Question 26

The correct answer is B.

Question 27

The correct answer is D.

Question 28

The correct answer is D.

Question 29

The correct answer is B.

Question 30

The correct answer is B.

Section B – Short-answer questions

Marks allocated are indicated by a number in square brackets, for example, [1] indicates that the line is worth one mark.

Question 1a

	Brief Description
Primary	The specific sequence of amino acids [1]
Secondary	Hydrogen bonding between carboxy (C=O) and amine (N-H) groups [1] of aligned amino acids forming alpha helices and beta sheets
Tertiary	Determines the overall 3D shape of the enzyme and is maintained by the interactions between the R-groups (includes hydrogen bonding, disulfide bridges, etc). [1] The catalytic activity of an enzyme often depends on its tertiary structure.

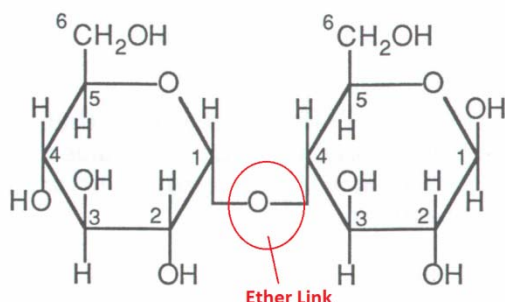
Question 1b

Experiment A: Heating an enzyme will increase the rate of reaction that the enzyme is catalysing but after 40° most enzymes in the human body will denature and no longer be functional. [1] Temperature affects the tertiary and secondary structure by breaking bonds between amino acids. [1]

Experiment B: Each enzyme has its own optimum pH but most enzymes denature [1] at very low pHs. The shape of their active site (tertiary structure) changes and they can no longer be functional. [1]

Question 2a

Maltose is made from a condensation reaction between 2 glucose molecules:



Question 2b

H₂O – disaccharides are formed through condensation reactions so water has to be a product.

Question 2c

$$M(C_6H_{12}O_6) = 180\text{g/mol}$$

When 9 molecules come together, 8 molecules of water are produced.

$$M(\text{polysaccharide}) = (9 \times 180) - (8 \times 18) = 1476\text{g/mol}$$

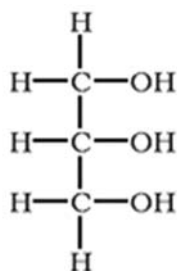
$$m(\text{polysaccharide}) = 1476 \times 0.12 = 177.12\text{g} = 177\text{g} \text{ (correct to 3 sig figs)}$$

Question 3a

Water is a polar. Fatty acids have a polar head (the carboxyl group) but a long non polar chain. This means that majority of the fatty acid molecule is hydrophobic. So it is unlikely that the long fatty acids will be soluble.

Question 3b

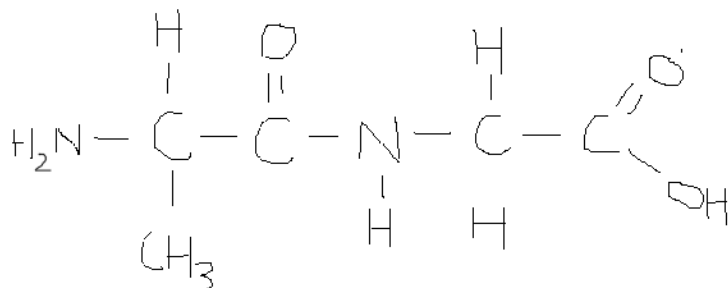
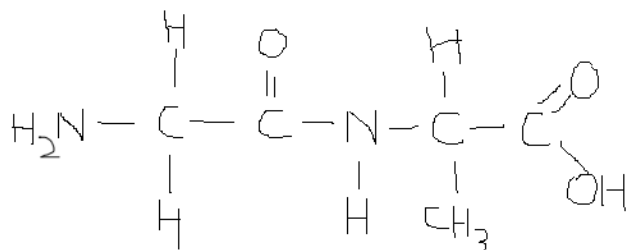
Glycerol

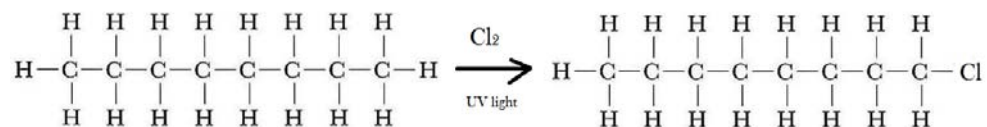
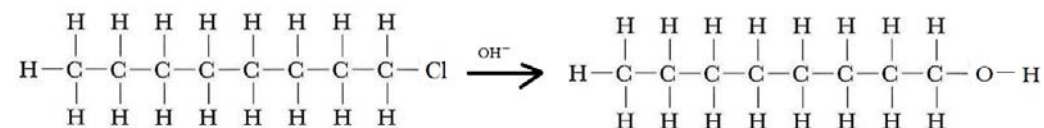
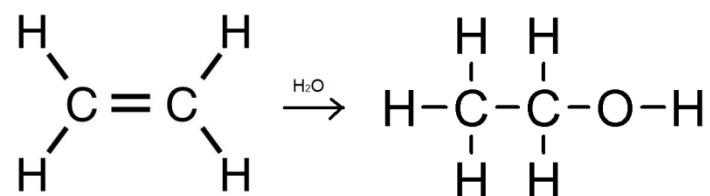
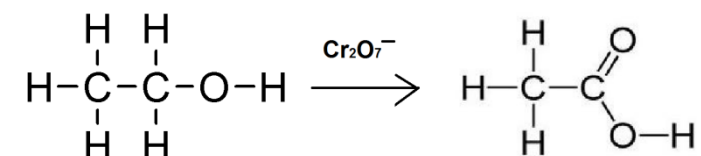
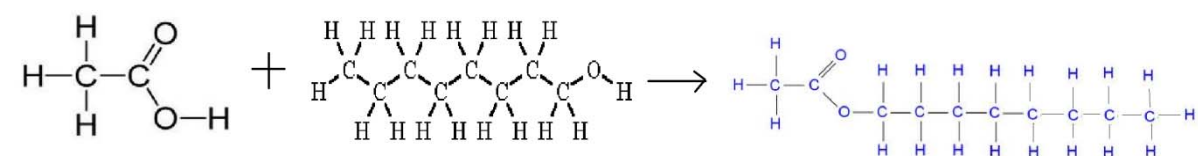
**Question 4a**

In a high pH environment amino acids act as acids. They do so by donating a hydrogen ion and developing an overall negative charge.

Question 4b

Zwitterion

Question 4c

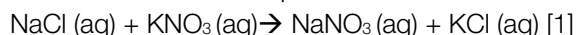
Question 5aOctane + $\text{Cl}_2 \rightarrow$ 1-chloro-octane1-chloro-octane + $\text{OH}^- \rightarrow$ 1-OctanolEthene + $\text{H}_2\text{O} \rightarrow$ EthanolEthanol + $\text{Cr}_2\text{O}_7^- \rightarrow$ Ethanoic acidEthanoic acid + octanol $\rightarrow$ octyl ethanoate**Question 5b**

Octyl Ethanoate

Question 6aIf the biscuit is 37% water then it has to $100-37=63\%$ biscuit. 63% of $5.6\text{g} = 3.528\text{g} = 3.5\text{g}$ (correct to 2 significant figures)**Question 6b**Table salt has the chemical formula NaCl . The metal element will therefore be sodium.

Question 6c

The reaction that takes place between NaCl and KNO₃ is as follows:

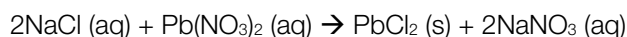


The student has been successful in making a new compound of sodium but sodium nitrate is aqueous.

This means that he will be unable to measure the amount of sodium as it does not precipitate. Therefore he will not be successful in calculating the amount of salt. [1]

Question 6d

Chlorine

Question 6e**Question 6f**

$$m(\text{PbCl}_2) = 1.3\text{g}$$

$$M(\text{PbCl}_2) = 207.2 + 35.5 + 35.5 = 278.2$$

$$n(\text{PbCl}_2) = 1.3/278.2 = 0.0047 \text{ mol (correct to 2 significant figures)} \quad [1]$$

$$n(\text{NaCl}) = 2 \times n(\text{PbCl}_2) = 2 \times 0.00472897 = 0.0093 \text{ mol (correct to 2 sf)} \quad [1]$$

$$m(\text{NaCl}) = 0.0094579439 \times (35.5 + 23) = 0.5467 = 0.55\text{g (correct to 2 sf)} \quad [1]$$

$$m(\text{biscuit}) = 3.528\text{g (from 7a)}$$

$$\% \text{ salt} = (0.5467/3.528) \times 100 = 16.78\% = 17\% \text{ (correct to 2 sf)} \quad [1]$$

Question 6g

Other impurities in the biscuit solution were also precipitated.

If the mass of the precipitate is higher the mol amount of salt will also be higher which would lead to a larger mass of salt, so a larger percentage of salt in the biscuit. [1]

Question 7a

Sample	Mean mass of sample (g)	Mean Absorbance	Iron Concentration (mg/L)
Fat Tony's concrete	0.4080	0.2695	38.83
Concrete Factory 1	0.4066	0.2002	28.63
Concrete Factory 2	0.4338	0.2730	39.35

1 mark is awarded for calculating the means, and 1 mark is awarded for calculating the concentrations.

Question 7b

Concrete Factory 2

Question 7c

$$c(\text{Fe in CF1}) = 28.63 \text{ mg/L}$$

$$m(\text{Fe}) = 28.63 \times 0.100\text{mL} = 2.863\text{mg} = 0.002863\text{g} \quad [1]$$

$$m(\text{original sample of concrete}) = 0.4066\text{g}$$

$$\%(\text{Fe in concrete}) = (0.002863/0.4066) \times 100 = 0.704\% \quad [1]$$

Question 7d

The iron cathode lamp was used so that the photons of light emitted onto the concrete sample corresponded to the specific quanta of energy that could be absorbed by iron atoms to promote electrons to an excited state. [1]

After excitation, the electrons return from an excited state to their ground state. When this occurs, light energy of a specific photon (in the case of iron, corresponding to an orange colour) is emitted. [1]

Question 8a

O₂ was added to the container. [1]

Question 8b

A = SO₃

B = SO₂ [1]

Students must identify both A and B to obtain the full mark.

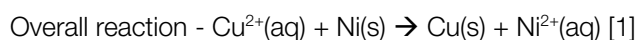
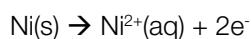
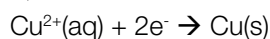
Question 8c

In accordance with Le Chatelier's Principle, as O₂ is added, the position of equilibrium moves to the right (there is a net forward reaction), so the amount and concentration of SO₃ (or A) increases [1] whereas the amount and concentration of SO₂ (or B) decreases. [1]

The concentration of O₂ decreases between 60 and 90 minutes but there is still a net increase in concentration of O₂ than its original concentration at 0 minutes. [1]

Question 8d

Between 60 and 90 minutes, there is a net forward reaction. The forward reaction is exothermic, [1] so the temperature of the mixture would increase. [1]

Question 9a

Cell potential (E°) = E° cathode – E° anode

$$= 0.34 - (-0.23)$$

$$= 0.57 \quad [1]$$

Question 9b

$$m(\text{Cu}) = 0.395\text{g}$$

$$n(\text{Cu}) = 0.395/63.5 = 6.22 \times 10^{-3}\text{mol}$$

$$n(\text{Ni}) \text{ lost} = n(\text{Cu}) \text{ formed} = 6.22 \times 10^{-3}\text{mol} \quad [1]$$

$$m(\text{Ni}) \text{ lost} = 6.22 \times 10^{-3} \times 58.7 = 0.365\text{g} \quad [1]$$

$$m(\text{Ni}) \text{ final} = 8.34\text{g} - 0.365 = 7.97\text{g} \quad [1]$$

Question 9c

$$\text{Initial amount of Ni}^{2+} = 0.100 \times 0.100 = 0.0100\text{mol}$$

$$n(\text{Ni}^{2+}) \text{ formed} = 6.22 \times 10^{-3}\text{mol}$$

$$n(\text{Ni}^{2+}) \text{ final} = 0.0100 + (6.22 \times 10^{-3}) = 0.01622 \text{ mol} \quad [1]$$

$$\text{final concentration} = 0.01622/0.100 = 0.162\text{M} \quad [1]$$

Question 10a

Le Chatelier's principle states: if an equilibrium system is subjected to an external change the system will adjust itself to partially oppose the effect of the change. [2]

Question 10b

Step 1 – this is a very exothermic reaction, so low temperatures would shift equilibrium to the right. However this reaction is carried out at a higher temperature (900°C – rather than low) to ensure an acceptable reaction rate. To compensate for the higher temperature used, the reaction should be carried out at low pressure, as the product side has fewer molecules than the reactant side. An appropriate catalyst would also increase rate of reaction and should therefore be used. Removing NO(g) as it forms would also push equilibrium to the right.

Step 2 – this reaction is not as exothermic as step 1 so a lower temperature will not have a large impact on yield, though it will still push equilibrium to the right. High pressure would also favour the forward reaction as the product side has fewer molecules. A suitable catalyst should be used and NO₂ should be continuously removed to keep the reaction proceeding in the forward direction.

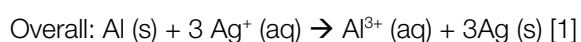
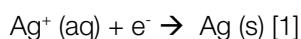
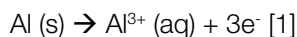
Step 3 – this is not an equilibrium reaction, so changes in temperature and pressure will not affect yield, despite the reaction being exothermic. However, increase in temperature will increase the rate of reaction.

Students were awarded [2] for each separate point they made for each step, as a number of points could be made.

Question 11a

Aluminium is a more reactive metal than silver. When both metals are in contact in an electrolyte, the solution forms a galvanic cell. [1]

The tarnish on the cutlery is a silver compound so it is converted back to solid silver according to the following equations:

**Question 11b**

This method does not wear away the silver surface [1], as other polishing methods do. Additionally, the materials needed are readily available in a domestic environment (sodium hydrogen carbonate (sodium bicarbonate) is baking soda).

Question 12a

$$K = [\text{OH}^-][\text{H}_3\text{BO}_3]/[\text{B}(\text{OH})_4^-] [1]$$

Question 12b

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$[\text{H}^+] = 10^{-\text{pH}} = 10^{-11.11} = 7.76 \times 10^{-12} \text{ M} [1]$$

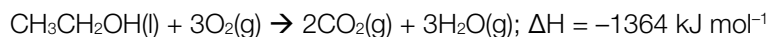
$$[\text{OH}^-] = 10^{-14}/10^{-11.11} = 1.29 \times 10^{-3} \text{ M} [1]$$

Question 12c

$$[\text{H}_3\text{BO}_3] = [\text{OH}^-] = 1.29 \times 10^{-3} \text{ M} [1]$$

Question 12d

$$K_a = 7.76 \times 10^{-12} \times 0.100 / 1.29 \times 10^{-3} = 6.01 \times 10^{-10} \text{ M [1]}$$

Question 13a

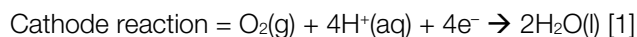
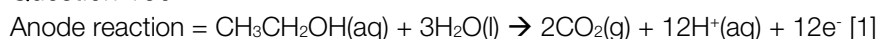
[1] for the balanced equation with states, [1] for the change in enthalpy value

Question 13b

$$\text{Total energy needed} = 8 \times 200 \times 10^3 \text{ kJ} = 1.6 \times 10^6 \text{ kJ [1]}$$

$$n(\text{CH}_3\text{CH}_2\text{OH})_{\text{required}} = 1.6 \times 10^6 / 1364 = 1173.02 \text{ mol [1]}$$

$$m(\text{CH}_3\text{CH}_2\text{OH})_{\text{required}} = 1173.02 \times 46 = 53958.94 \text{ g} = 54.0 \text{ kg [1]}$$

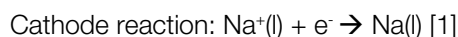
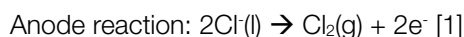
Question 13c**Question 13d**

$$Q = n(\text{e}^-) \times F = 12 \times 96500 = 1158000 \text{ [1]}$$

$$\text{Energy} = V \times Q = 1.15 \times 1158000 = 1.33 \times 10^6 \text{ J [1]}$$

Question 13e

A fuel cell is more efficient at producing energy [1]. It also produces less pollution.

Question 14a**Question 14b**

The iron mesh prevents chlorine gas formed at the anode coming into contact with sodium formed at the cathode. If the two were allowed to come into contact, they would spontaneously react to reform NaCl. [1]

Question 14c

$$Q = t \times I = 1.00 \times 10^4 \times 96.5 = 9.65 \times 10^5 \text{ C [1]}$$

$$n(\text{e}^-) = 9.65 \times 10^5 / 96500 = 10$$

$$n(\text{Cl}_2) = \frac{1}{2} n(\text{e}^-) = 5 \text{ mol [1]}$$

$$V(\text{Cl}_2) = 5 \times 22.4 = 112 \text{ L [1]}$$